



Province of the
EASTERN CAPE
EDUCATION

**NATIONAL
SENIOR CERTIFICATE**

GRADE 12

JUNE 2019

**PHYSICAL SCIENCES P2
(CHEMISTRY)**

MARKS: 150

TIME: 3 hours



This question paper consists of 16 pages, including a formula sheet
and data sheets.

INSTRUCTIONS AND INFORMATION

1. Write your full NAME and SURNAME in the appropriate spaces on the ANSWER BOOK.
2. This question paper consists of SEVEN questions.
3. Start EACH question on a NEW page in the ANSWER BOOK.
4. Number the answers correctly according to the numbering system used in this question paper.
5. Leave ONE line between two sub-questions, for example between QUESTION 2.1 and QUESTION 2.2.
6. You may use a non-programmable calculator.
7. You may use appropriate mathematical instruments.
8. You are advised to use the attached DATA SHEETS.
9. Show ALL formulae and substitutions in ALL calculations.
10. Round off your FINAL numerical answers to a minimum of TWO decimal places.
11. Give brief motivations, discussions, et cetera where required.
12. Write neatly and legibly.

QUESTION 1: MULTIPLE-CHOICE QUESTIONS

Four options are provided as possible answers to the following questions. Each question has only ONE correct answer. Write only the letter (A–D), corresponding to the correct answer of your choice, next to the question number (1.1–1.10) in the ANSWER BOOK, for example 1.11 D.

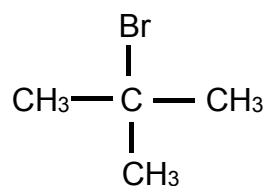
1.1 Which ONE of the following compounds has a DOUBLE BOND in its structure?

- A Ethane
 - B Ethene
 - C Polyethene
 - D Bromoethane
- (2)

1.2 By definition the boiling point of a liquid is the temperature at which ...

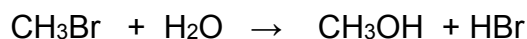
- A a liquid changes to vapour.
 - B the vapour pressure equals atmospheric pressure.
 - C the vapour pressure is less than the atmospheric pressure.
 - D the vapour pressure is greater than the atmospheric pressure.
- (2)

1.3 Which ONE of the following is the CORRECT IUPAC name of the compound shown below?



- A 2-bromobutane
 - B 2-methyl-2-bromopropane
 - C 2-bromo-2-methylpropane
 - D 2,2,2-trimethylbromomethane
- (2)

1.4 Consider the reaction represented by the equation below:



The TYPE of reaction shown above is ...

- A substitution.
 - B combustion.
 - C esterification.
 - D polymerisation.
- (2)

- 1.5 The boiling points of propanal and propanone are compared.

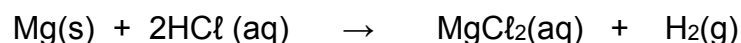
Which ONE of the following is the INDEPENDENT variable in this comparison?

- A Molar mass
 - B Chain length
 - C Hydrogen bonds
 - D Homologous series
- (2)

- 1.6 Which ONE of the following factors will change the value of the equilibrium constant, K_c ?

- A Pressure
 - B Surface area
 - C Temperature
 - D Concentration
- (2)

- 1.7 The reaction of magnesium powder and hydrochloric acid is used to investigate the factors that affect reaction rate. The balanced equation for the reaction is:



In ALL the experiments the hydrochloric acid is in EXCESS and the magnesium powder is completely covered.

The results of the experiments are shown in the table below:

Experiment	Concentration of HCl (mol.dm ⁻³)	Volume of HCl (cm ³)	Mass of Mg (g)
P	1,5	200	1,8
Q	1,2	400	1,8
R	0,8	200	1,4
S	1,5	200	1,5

Which experiments will produce the SAME amount of hydrogen at same temperature and pressure?

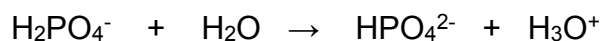
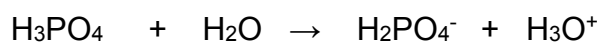
- A P and Q
 - B Q and R
 - C R and S
 - D P and S
- (2)

- 1.8 Diluted hydrochloric acid is spilt on floor tiles in the laboratory.

Which ONE of the following substances can be used to neutralise the acid?

- A NaCl
 - B NH₄Cl
 - C H₂SO₄
 - D Na₂CO₃
- (2)

1.9 Study the ionisation reactions given below:

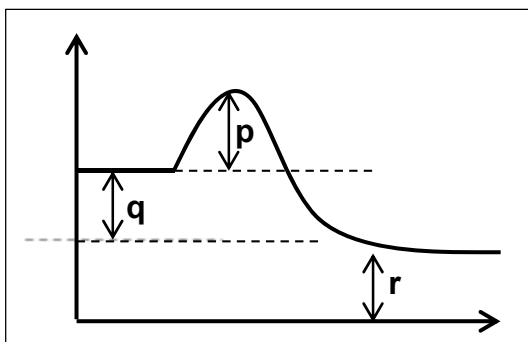
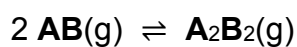


In the reactions H_2O acts as a(n) ...

- A base.
- B ampholyte.
- C weak acid.
- D strong acid.

(2)

1.10 The potential energy diagram shown below is for the reversible hypothetical reaction shown below:



Consider the following statements about the reaction:

- I** The forward reaction is endothermic
- II** The value of ΔH for the REVERSE reaction is equal to **q**
- III** The activation energy for the reverse reaction is equal to **p + q**

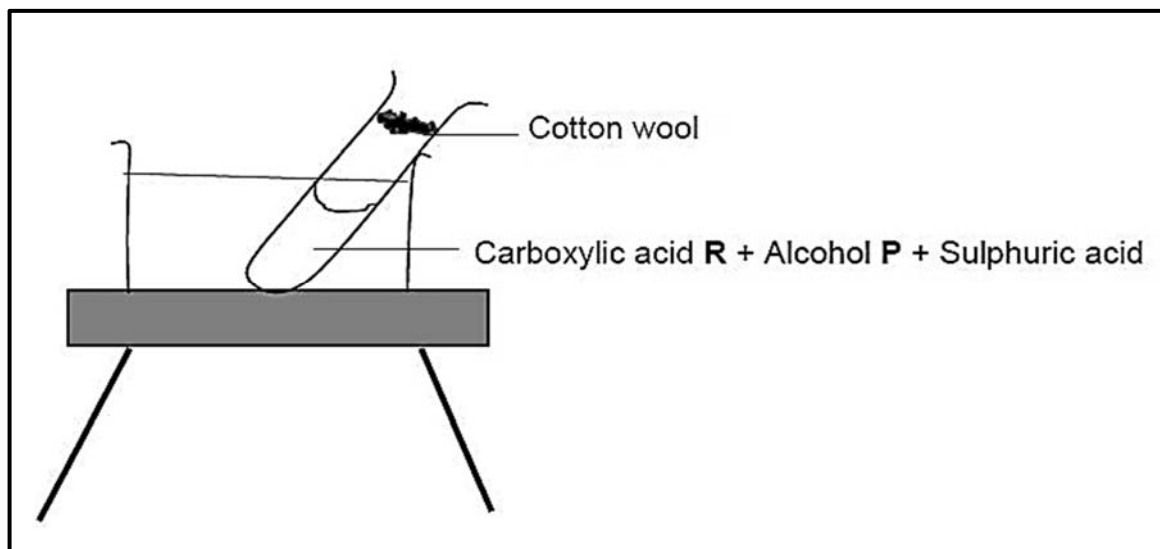
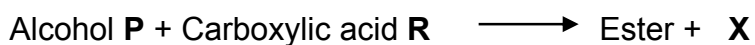
Which of the above statements is/are TRUE?

- A **I** only
- B **I** and **II**
- C **III** only
- D **II** and **III**

(2)
[20]

QUESTION 2

A test tube containing an alcohol **P**, carboxylic acid **R** and sulphuric acid is heated in a water bath, as illustrated below. A piece of wet cotton wool is placed at the mouth of the test tube.



- 2.1 Is the sulphuric acid that is used CONCENTRATED or DILUTE? (1)
- 2.2 What is the purpose of placing cotton wool at the mouth of the test tube? (1)
- 2.3 Write down the molecular formula of the inorganic product, **X**. (1)
- 2.4 Alcohol **P** is a positional isomer of propan-2-ol.
Write down the IUPAC name of alcohol **P**. (2)
- 2.5 Chemical analysis shows that carboxylic acid **R** contains 54,55% C, 9,1% H and **x** % O.

The molar mass of the **ester** produced is $130 \text{ g} \cdot \text{mol}^{-1}$.

Determine by calculation the MOLECULAR formula of carboxylic acid **R**. (8)

[13]

QUESTION 3

- 3.1 The boiling points of three compounds (**A**, **B** and **C**) are compared in the table below:

	COMPOUND	BOILING POINT (°C)
A	Pentan-1-ol	138
B	Butanoic acid	164
C	Methyl propanoate	68

- 3.1.1 For compound **C** write down the:

- (a) Name of the homologous series to which it belongs (1)
- (b) STRUCTURAL formula (2)

- 3.1.2 Give a reason why this is a fair comparison. (2)

- 3.1.3 Explain why hydrogen bonds are stronger in compound **B** than in compound **A**. (2)

- 3.2 The vapour pressure values of a secondary and a tertiary alcohol are compared in the table below:

	COMPOUND	MOLECULAR FORMULA	VAPOUR PRESSURE (kPa at 20 °C)
D	Secondary alcohol	C ₄ H ₁₀ O	X
E	Tertiary alcohol	C ₄ H ₁₀ O	Y

- 3.2.1 Are compounds **D** and **E** STRUCTURAL isomers?
Write down Yes or No.

- Give a reason for the answer. (3)

- 3.2.2 Which ONE of the two values, **X** or **Y** is HIGHER?

- Fully explain the answer. (4)

[14]

QUESTION 4

- 4.1 Study the two reactions shown below. **P** is an inorganic compound.
R is an organic compound.



- 4.1.1 Write down the TYPE of reaction represented by:

- (a) Reaction **I** (1)
(b) Reaction **II** (1)

- 4.1.2 Write down the:

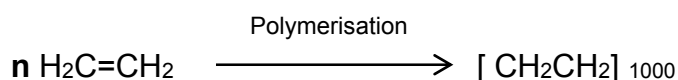
- (a) STRUCTURAL formula of 1,2-dichloro-3-methylbutane (3)
(b) Name of the inorganic reactant **P** (1)

- 4.1.3 For reaction **II** write down:

- (a) The structural formula and IUPAC name of the major product, **R** (4)
(b) ONE reaction condition (1)

- 4.2 The equation shown below shows **n** molecules of H₂C=CH₂ reacting to produce a polymer.

H₂C=CH₂ in the equation below represents small organic molecules that can be covalently bonded to each other in a repeating pattern

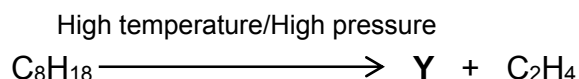


- 4.2.1 Is this reaction an example of ADDITION or CONDENSATION polymerisation? (1)

Write down:

- 4.2.2 The value of **n** (1)
4.2.3 A term used for the underlined phrase (1)
4.2.4 ONE use of the polymer produced in this reaction (1)

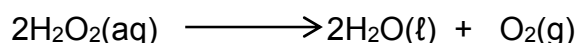
- 4.3 The following equation represents the cracking reaction of octane (C_8H_{18}).
Y is a straight chain organic compound.



- 4.3.1 Define the term *cracking* reaction. (2)
- 4.3.2 Is this reaction an example of THERMAL CRACKING or CATALYTIC CRACKING? (1)
- 4.3.3 Write down the STRUCTURAL FORMULA and IUPAC name of compound Y. (4)
- [22]**

QUESTION 5

Hydrogen peroxide (H_2O_2) decomposes according to the following equation:



- 5.1 State THREE factors that can INCREASE the rate of this reaction. (3)
- 5.2 Define the term *reaction rate*. (2)

The rate of decomposition of hydrogen peroxide is investigated in THREE experiments. The initial temperature of H_2O_2 is the same in ALL the experiments.

The reaction runs to completion in ALL the experiments.

- 5.3 In **experiment 1** a solution of hydrogen peroxide is heated to $35^\circ C$.

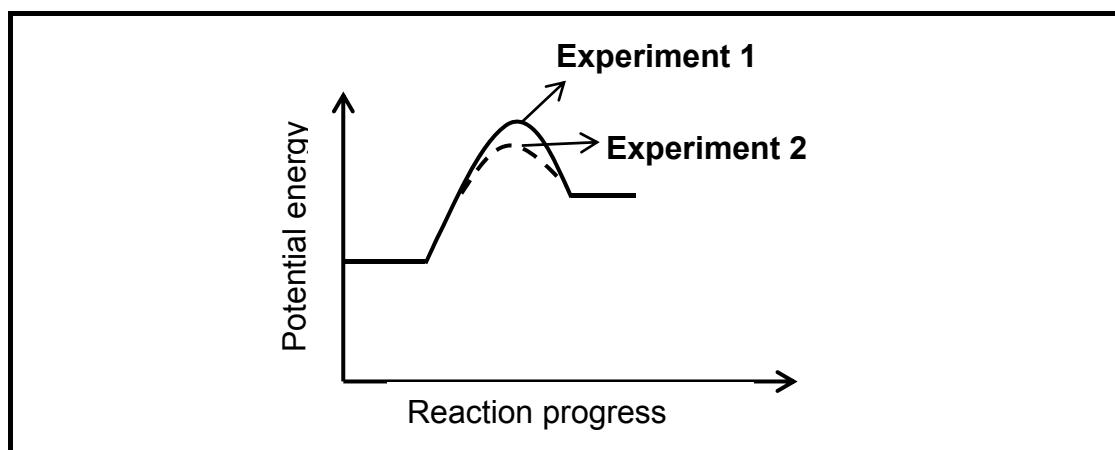
The concentration of H_2O_2 was measured at different time intervals during the experiment.

The following results were obtained:

TIME (MINUTES)	H_2O_2 CONCENTRATION ($mol \cdot dm^{-3}$)
0	1,9
15	1,45
55	1,10
100	0,85
215	0,60

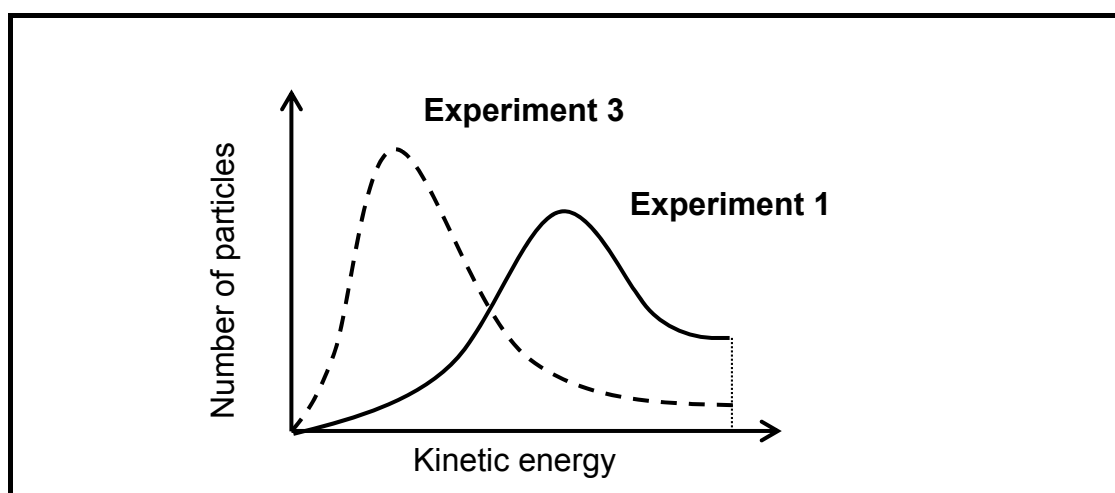
- 5.3.1 Calculate the average reaction rate, in $mol \cdot dm^{-3} \cdot min^{-1}$ during the first 15 minutes. (3)
- 5.3.2 Give a reason why the average reaction rate is HIGHER during the first 15 minutes, compared to the time interval 100 to 215 seconds. (2)

- 5.4 In **experiment 2**, a small amount of manganese dioxide is added to the same H_2O_2 solution (used in **experiment 1**) and the mixture is heated to 35°C .



- 5.4.1 Is the reaction EXOTHERMIC or ENDOTHERMIC? (1)
- 5.4.2 Use the collision theory to explain the effect of manganese dioxide on the rate of decomposition of H_2O_2 . (4)
- 5.5 In **experiment 3** the same volume of the solution of H_2O_2 (used in **experiment 1**) is used but ONLY the temperature to which the solution is heated is changed.

The Maxwell-Boltzmann distribution curves for **experiment 1** and **3** are shown below:



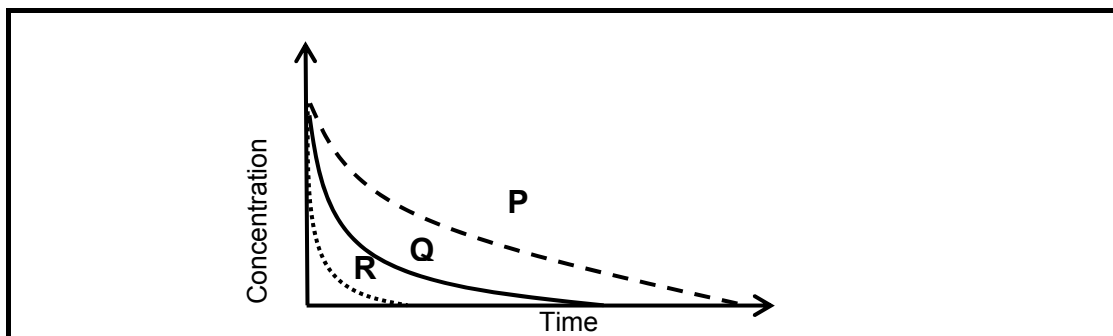
- 5.5.1 Which experiment (1 or 3) was carried out at a HIGHER temperature?
Give a reason for the answer. (2)
- 5.5.2 How does the area under the graph of **experiment 1** compare to the area under the graph of **experiment 3**?
Write down GREATER THAN, EQUAL TO or LESS THAN.
Explain the answer. (2)

5.6 In **experiment 2** the total volume of oxygen produced at 35 °C is 0,2 dm³.

Calculate the mass of H₂O₂ that decomposed in **experiment 2** if the molar volume of oxygen at 35 °C is 24,8 dm³·mol⁻¹.

(5)

5.7 The graphs below show how the concentration of H₂O₂ changes with time in the experiments.



Which graph (**P**, **Q** or **R**) represents the results of:

5.7.1 Experiment 1

(1)

5.7.2 Experiment 2

(1)

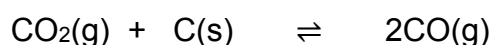
5.7.3 Experiment 3

(1)

[27]

QUESTION 6

Carbon dioxide gas, $\text{CO}_2(\text{g})$, reacts with carbon, $\text{C}(\text{s})$, according to the following balanced equation:



The system reaches equilibrium at $1\,000\text{ }^\circ\text{C}$.

6.1 Define the term *chemical equilibrium*. (2)

6.2 How will EACH of the following changes affect the number of moles of $\text{CO}(\text{g})$ at equilibrium?

Choose from INCREASES, DECREASES or REMAINS CONSTANT.

6.2.1 The addition of more carbon, C (1)

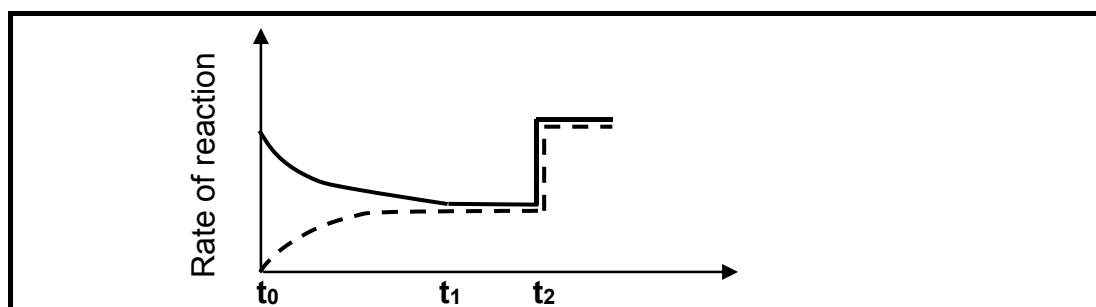
6.2.2 The addition of more CO_2 (1)

6.2.3 Decreasing the pressure by increasing the volume (1)

6.3 Use Le Chatelier's principle to explain the answer to QUESTION 6.2.3 above. (3)

6.4 The sketch graph given below shows the changes in the rates of reaction from the moment the reagents are introduced into the container. The reaction reaches equilibrium for the first time at t_1 .

A change is made to the equilibrium mixture at t_2 .



6.4.1 Which reaction (FORWARD or REVERSE) is represented by the broken line? (1)

6.4.2 What change was made at t_2 ? (2)

Initially 104,72 g of CO_2 was heated with EXCESS carbon C in a sealed flask of volume 250 cm^3 . The equilibrium mixture was found to contain 1,90 mol of CO_2 .

6.5 Calculate the value of the equilibrium constant, K_c at $1\,000\text{ }^\circ\text{C}$. (8)

- 6.6 The value of K_c for the reaction decreases to 0,333 when the temperature is changed to 25 °C.

6.6.1 Is there a HIGH or LOW yield at 25 °C?

Give a reason for the answer. (2)

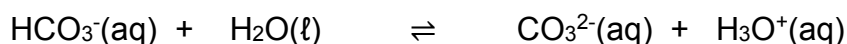
6.6.2 Is the FORWARD reaction ENDOTHERMIC or EXOTHERMIC?

Explain the answer fully. (4)

[25]

QUESTION 7

- 7.1 The hydrogen carbonate ion (HCO_3^-) undergoes hydrolysis according to the equation:



7.1.1 Define the term *hydrolysis*. (2)

7.1.2 Give a reason why HCO_3^- is regarded as an ampholyte. (2)

Write down the FORMULA of the:

7.1.3 Conjugate acid of CO_3^{2-} (1)

7.1.4 Conjugate acid of HCO_3^- (not in the equation) (1)

- 7.2 A weak monoprotic acid is dissolved in water to produce a solution of accurately known concentration.

7.2.1 Write down a term for the underlined phrase. (1)

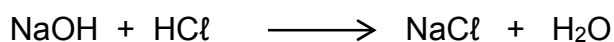
7.2.2 Which ONE of the following is an example of a *weak monoprotic acid*?

CH_3COOH	HCl	H_2CO_3
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The weak monoprotic acid has a concentration of 0,2 mol·dm⁻³. (2)

7.2.3 Calculate the volume to which 20 cm³ of the acid must be diluted to produce a solution with a concentration of 0,16 mol·dm⁻³. (3)

- 7.3 During a titration a learner adds 30 cm³ of NaOH solution of concentration 0,1 mol.dm⁻³ to an Erlenmeyer flask and titrates this solution with 25 cm³ HCl solution. The balanced equation for the reaction that takes place is:



- 7.3.1 Name this type of reaction. (1)

- 7.3.2 At what pH range will a suitable indicator for this titration change colour?

Choose from:

6,8 to 7,2	3 to 5	8,6 to 10
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- Give a reason for the answer. (3)

- 7.3.3 Calculate the concentration of the HCl that was used to neutralise the NaOH in the Erlenmeyer flask. (4)

The learner passes the endpoint by adding an extra 8 cm³ of the HCl solution.

- 7.3.4 Define the term *endpoint*. (2)

- 7.3.5 Calculate the pH of the solution in the flask after the addition of the extra 8 cm³ of HCl. (7)

[29]

TOTAL: 150

**NATIONAL SENIOR CERTIFICATE
NASIONALE SENIOR SERTIFIKAAT**

**DATA FOR PHYSICAL SCIENCES GRADE 12
PAPER 2 (CHEMISTRY)**

**GEGEWENS VIR FISIIESE WETENSKAPPE GRAAD 12
VRAESTEL 2 (CHEMIE)**

TABLE 1: PHYSICAL CONSTANTS/TABEL 1: FISIIESE KONSTANTES

NAAM/NAME	SIMBOOL/SYMBOL	WAARDE/VALUE
Standard pressure <i>Standaarddruk</i>	p^θ	$1,013 \times 10^5 \text{ Pa}$
Molar gas volume at STP <i>Molêre gasvolume teen STD</i>	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature <i>Standaardtemperatuur</i>	T^θ	273 K
Charge on electron <i>Lading op elektron</i>	e	$-1,6 \times 10^{-19} \text{ C}$
Avogadro's constant <i>Avogadro se konstante</i>	N_A	$6,02 \times 10^{23} \text{ mol}^{-1}$

TABLE 2: FORMULAE/TABEL 2: FORMULES

$n = \frac{m}{M} \text{ or/of}$ $n = \frac{N}{N_A} \text{ or/of}$ $n = \frac{V}{V_m}$	$c = \frac{n}{V} \text{ or/of } c = \frac{m}{MV}$ $\frac{c_a V_a}{c_b V_b} = \frac{n_a}{n_b}$	$\text{pH} = -\log[\text{H}_3\text{O}^+]$ $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14}$ at /by 298K
$E^\theta_{\text{cell}} = E^\theta_{\text{cathode}} - E^\theta_{\text{anode}} / E^\theta_{\text{sel}} = E^\theta_{\text{katode}} - E^\theta_{\text{anode}}$ $E^\theta_{\text{cell}} = E^\theta_{\text{reduction}} - E^\theta_{\text{oxidation}} / E^\theta_{\text{sel}} = E^\theta_{\text{reduksie}} - E^\theta_{\text{oksidasie}}$ $E^\theta_{\text{cell}} = E^\theta_{\text{oxidising agent}} - E^\theta_{\text{reducing agent}} / E^\theta_{\text{sel}} = E^\theta_{\text{oksideermiddel}} - E^\theta_{\text{reduseermiddel}}$		

TABLE 3: THE PERIODIC TABLE OF ELEMENTS/TABEL 3: DIE PERIODIEKE TABEL VAN ELEMENTE

KEY/ SLEUTEL																	
Atomgetal Atomic number																	
1 (I)	2 (II)	3	4	5	6	7	8	9	10	11	12	13 (III)	14 (IV)	15 (V)	16 (VI)	17 (VII)	18 (VIII)
1 H 1																	2 He 4
3 Li 7	4 Be 9															9 F 19	10 Ne 20
11 Na 23	12 Mg 24															17 Cl 35,5	18 Ar 40
19 K 39	20 Ca 40	21 Sc 45	22 Ti 48	23 V 51	24 Cr 52	25 Mn 55	26 Fe 56	27 Co 59	28 Ni 59	29 Cu 63,5	30 Zn 65	31 Ga 70	32 Ge 73	33 As 75	34 Se 79	35 Br 80	36 Kr 84
37 Rb 86	38 Sr 88	39 Y 89	40 Zr 91	41 Nb 92	42 Mo 96	43 Tc 98	44 Ru 101	45 Rh 103	46 Pd 106	47 Ag 108	48 Cd 112	49 In 115	50 Sn 119	51 Sb 122	52 Te 128	53 I 127	54 Xe 131
55 Cs 133	56 Ba 137	57 La 139	72 Hf 179	73 Ta 181	74 W 184	75 Re 186	76 Os 190	77 Ir 192	78 Pt 195	79 Au 197	80 Hg 201	81 Tl 204	82 Pb 207	83 Bi 209	84 Po 210	85 At 210	86 Rn 222
87 Fr 227	88 Ra 226	89 Ac															

Elektronegatiwiteit
Electronegativity

Benaderde relatiewe atoommassa
Approximate relative atomic mass

Simbool
Symbol

29
Cu
63,5

58 Ce 140	59 Pr 141	60 Nd 144	61 Pm	62 Sm 150	63 Eu 152	64 Gd 157	65 Tb 159	66 Dy 163	67 Ho 165	68 Er 167	69 Tm 169	70 Yb 173	71 Lu 175
90 Th 232	91 Pa	92 U 238	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr